

## Mixed Stoichiometry Review Answers

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### Mixed Stoichiometry Review Answers

Stoichiometry Review Answers 1. a.  $\text{Na}_3\text{PO}_4$  b.  $\text{Ca}(\text{NO}_3)_2$  Na = 3 mol  $\times$  22.99 g/mol = 68.97 g Ca = 1 mol  $\times$  40.08 g/mol = 40.08 g P = 1 mol  $\times$  30.97 g/mol = 30.97 g N = 2 mol  $\times$  14.01 g/mol = 28.02 g O = 4 mol  $\times$  16.00 g/mol = 64.00 g O = 6 mol  $\times$  16.00 g/mol = 96.00 g 163.94 g 164.10 g c.  $\text{Ca}_3(\text{PO}_4)_2$  d.

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Modern Chemistry 2 Stoichiometry CHAPTER 9 REVIEW Stoichiometry MIXED REVIEW SHORT ANSWER Answer the following questions in the space provided. 1. Given the following equation:  $\text{C}_3\text{H}_4(\text{g}) + x\text{O}_2(\text{g}) \rightarrow 3\text{CO}_2(\text{g}) + 2\text{H}_2\text{O}(\text{g})$  \_\_\_\_ a. What is the value of the coefficient x in this equation?

### Modern Chemistry Chapter 9 Stoichiometry Mixed Review Answers

ANSWER KEY. Mixed Stoichiometry Problems . 1.  $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$ . a). How many moles of  $\text{H}_2$  would be required to produce 5.0 moles of water? 5.0 moles water. b). What mass of  $\text{H}_2\text{O}$  is formed when  $\text{H}_2$  reacts with 384 g of  $\text{O}_2$ ? 432g  $\text{H}_2$ . 2.  $\text{H}_2\text{SO}_4 + 2\text{NaOH} \rightarrow \text{Na}_2\text{SO}_4 + 2\text{H}_2\text{O}$ . a). Balance this equation. Look above. b).

### Mixed Stoichiometry Problems

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### Modern Chemistry Chapter 9 Mixed Review Stoichiometry Answers

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### Stoichiometry Mixed Problems

Chemistry Worksheet on Stoichiometry Mixed Review. Assume all reactions go to completion. Write the formula equation, balance the equations, and solve the problems. Draw a rectangle around the answer and don't forget the units. Methane ( $\text{CH}_4$ ) combines with oxygen to form carbon dioxide and water. Balanced equation:

### Chemistry Worksheet on Stoichiometry Mixed Review

Mixed Stoichiometry Practice. Potassium Chlorate decomposes into potassium chloride and oxygen gas. Balanced Equation:  $\text{KClO}_4 \rightarrow \text{KCl} + \text{O}_2$ . How many grams of oxygen are produced when 3.0 moles of potassium chlorate decompose completely? Butane ( $\text{C}_4\text{H}_{10}$ ) undergoes combustion.

### Mixed Stoichiometry Practice

CHAPTER 9 REVIEW Stoichiometry MIXED REVIEW SHORT ANSWER Answer the following questions in the space provided. 1. Given the following equation:  $\text{C}_3\text{H}_4(\text{g}) + x\text{O}_2(\text{g}) \rightarrow 3\text{CO}_2(\text{g}) + 2\text{H}_2\text{O}(\text{g})$  4 a. What is the value of the coefficient x in this equation? 40.07 g/mol b. What is the molar mass of  $\text{C}_3\text{H}_4$ ? 2 mol O 2:1 mol H 20 c. What is the mole ratio of O 2 to H

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MIXED REVIEW [P.E.R. REVIEW Stoichiometry SHORT ANSWER Answer the following questions in the space provided. 1. Given the following equation:  $\text{C}_3\text{H}_4(\text{g}) + x\text{O}_2(\text{g}) \rightarrow 3\text{CO}_2(\text{g}) + 2\text{H}_2\text{O}(\text{g})$  \_\_\_\_ a. What is the value of the coefficient x in this equation? 40.07 g/mol h. What is the molar mass of  $\text{C}_3\text{H}_4$ ? 2 mol 02:1 mol  $\text{H}_2\text{O}$  c.

### Date. FCHAPJ REVIEW.

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